

$$n(\text{H}_2) = \frac{3,36}{22,4} = 0,15 \text{ mol}$$

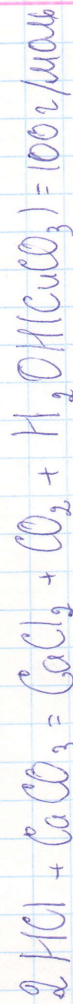
$$m(\text{Mg}) = 0,15 \cdot 24 = 3,624 \text{ gram}$$

$$m(\text{Cu}) \text{ dalam } = 25 - 36 = 21,42$$

$$m(\text{Mg}) = \frac{21,4}{25} = 0,856$$

$$\text{Omben} : 85,6\%$$

54



$$\frac{x}{100} = \frac{1,12}{22,420} = 52 \text{ CaCO}_3$$

$$5 - 100$$

$$x - 15$$

$$1 = 0,952$$

$$5 - 0,95 = 5,26$$

2015  
Date

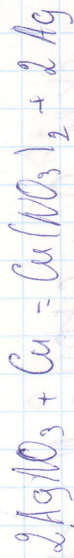
Perhitungan:

$$m(p-pa) = 1694 \text{ m}(\text{AgNO}_3) = w(\text{AgNO}_3) \cdot m(p-pa) =$$

$$= 0,025 \cdot 169 = 4,23 \text{ g}$$

$$W(\text{AgNO}_3) = 0,025$$

$$n(\text{Ag}) = ?$$



$$M(\text{AgNO}_3) = 108 + 14 + 16 \cdot 3 = 170 \text{ g/mol}$$

$$n(\text{Ag}) = x = 4,23 \cdot 2 : 12 \cdot 170 = 0,025 \text{ mol}$$

58!

230!